

C4

Question	Answer
What is relative atomic mass?	It is the mass of an atom (relative to C-12)
What is relative formula mass of a compound?	It is the sum of the relative atomic masses of the atoms in the formula
Give an example of relative formula mass for water.	H ₂ O H x 2 O x 1 Ar of H = 1, 1 x 2 = 2 Ar of O = 16, 16 x 1 = 16 16 + 2 = 18
Why does the mass of the products not always equal the mass of the reactants in an open system?	A gas may be produced which will go into the surroundings but it still has mass
What is meant by uncertainty of results?	That repeat measurements may be quite different
How can uncertainty be reduced?	Look at the range of a set of measurements about the mean, dispose of results not within a close enough range
What is the state symbol for a solution?	(aq) – aqueous
What is the unit of concentration?	g/dm ³ - Grams per dm ³

C4 HT

Question	Answer
What is a mole (mol)?	The mass of 1 mole of a substance is the relative atomic/formula mass in grams
What is Avogadro's constant?	The number of atoms, molecules or ions in a mole of a given substance
What does 2H ₂ O mean in an equation?	That there are 2 moles of water
What is the limiting reactant?	The reactant that limits the amount of products
How can you ensure that all of the limiting reactant is used up?	Use an excess of the other reactant in the reaction