

C2 – Period table		
1	What is the family name of group 0	Noble Gases
2	What is the family name of group 1	Alkali Metals
3	What is the family name of group 7	Halogens
4	What happens to the boiling point of group 0 as you go down the group?	The boiling points of the noble gases increase with increasing relative atomic mass (going down the group).
5	Describe the reactivity of group 0 elements and explain whether they form molecules.	They are unreactive and do not easily form molecules because their atoms have stable arrangements of electrons.
6	How many electrons are in the outer shell of group 0 elements	The noble gases have eight electrons in their outer shell, except for helium, which has only two electrons.
7	Describe the reactivity of group 1 elements	In Group 1, the reactivity of the elements increases going down the group.
8	How many electrons are in the outer shell of group 1 elements	They have a single electron in their outer shell.
9	What happens to the melting point and boiling point of group 1 as you go down the group?	As you go down the elements in group one their boiling points and melting points get lower
10	Give some of the properties of group 1	They have a low density and float on water, and they are soft. They also all react with water to form hydrogen gas and a metal hydroxide.
11	What type of compound is usually formed from Group 1 elements?	These elements react with non-metals creating ionic compounds. In which the metal ion has a charge of +1
12	Describe the reactivity of group 7 elements	In Group 7, the reactivity of the elements decreases going down the group. A more reactive halogen can displace a less reactive halogen from an aqueous solution of its salt.
13	How many electrons are in the outer shell of group 7 elements	The elements in Group all have seven electrons in their outer shell.
14	What happens to the melting point and boiling point of group 7 as you go down the group?	In Group 7, the further down the group an element is the higher its relative molecular mass, melting point and boiling point.
15	What type of compound is usually formed from Group 7 elements?	They react with metals to form ionic compounds in which the halide ion carries a charge of -1 . They can also form simple covalent molecules when reacting with other non-metal elements.
17	Describe the reactions of group 1 metals with a) oxygen b) chlorine c) water	a) quickly forms an oxide layer b) forms an ionic bond and produces a salt c) reacts vigorously producing Hydrogen gas and a strong alkali (metal hydroxide)